GCSE (9-1)

Combined Chemistry 2

**Topics for Paper 2**

**Topic 6 – Groups in the periodic table Group 1**

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| **Students should:** | **Maths skills** |
| 6.1 | Explain why some elements can be classified as:alkali metals (group 1) – *they have 1 electron in the outside shell*halogens (group 7) – *they have 7 electrons in the outside shell*noble gases (group 0) – *they have full outer shells*based on their position in the periodic table |  |
| 6.2 | Recall that alkali metals: a are softb have relatively low melting points |  |
| 6.3  | Describe the reactions of lithium, sodium and potassium with water2Li + 2H2O 🡪 2LiOH + H22Na + 2H2O 🡪 2NaOH + H22K+ 2H2O 🡪 2KOH + H2 |  |
| 6.4 | Describe the pattern in reactivity of the alkali metals, lithium, sodium and potassium, with water; and use this pattern to predict the reactivity of other alkali metals* *Reactivity increases as you go down the group*
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| 6.5 | Explain this pattern in reactivity in terms of electronic configurations* *The more readily a metal loses its outer electron the more reactive it is.*
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**Group 7**

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| **Students should:** | **Maths skills** |
| 6.6 | Recall the colours and physical states of chlorine, bromine and iodine at room temperature* *Chlorine = green gas*
* *Bromine = red-brown liquid that gives off an orange vapour*
* *Iodine = dark grey solid that gives off a purple vapour*
 |  |
| 6.7 | Describe the pattern in the physical properties of the halogens, chlorine, bromine and iodine, and use this pattern to predict the physical properties of other halogens* Reactivity decreases as you go down the group.
* The melting point and boiling point increase as you go down the group.
 | 1d 2c |
| 6.8 | Describe the chemical test for chlorine* *Damp blue litmus paper turns white*
 |  |
| 6.9 | Describe the reactions of the halogens, chlorine, bromine and iodine, with metals to form metal halides, and use this pattern to predict the reactions of other halogens*Metal + Halogen 🡪 Metal halide**For example:*2Na + Cl2 🡪 2NaCl2Na + Br2 🡪 2NaBr2Na + I2 🡪 2NaI |  |
| 6.10 | Recall that the halogens, chlorine, bromine and iodine, form hydrogen halides which dissolve in water to form acidic solutions, and use this pattern to predict the reactions of other halogensCl2 + H2 🡪 2HCl |  |
| 6.11 | Describe the relative reactivity of the halogens chlorine, bromine and iodine, as shown by their displacement reactions with halide ions in aqueous solution, and use this pattern to predict the reactions of astatine* *A more reactive halogen will displace a less reactive one, for example:*

*Chlorine + Potassium Bromide 🡪 Bromine + Potassium chloride* |  |
| 6.12 | **Explain why these displacement reactions are redox reactions in terms of gain and loss of electrons, identifying which of the substances are oxidised and which are reduced.*** *The halogens gain electrons.*
* *Halide ions lost electrons*
* *For example;*

*Cl2 + 2Br- 🡪 Br2 + 2Cl-* |  |

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| **Students should:** | **Maths skills** |
| 6.13 Explain the relative reactivity of the halogens in terms of electronic configurations *- As you go down the group the element are less reactive because it is harder to attract the extra electron as the atomic radius is bigger.* |  |

**Group 0**

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| **Students should:** | **Maths skills** |
| 6.14 | Explain why the noble gases are chemically inert, compared with the other elements, in terms of their electronic configurations* *They have full outer shells so don’t gain or lose electrons*
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| 6.15 | Explain how the uses of noble gases depend on their inertness, low density and/or non-flammability* *Argon is used in filament lamps as it’s non flammable. It stops the filament from burning away.*
* *Argon and helium can be used to protect metals that are being welded. It stops them reacting with oxygen.*
* *Helium is used in airships and party balloons. Helium has a lower density than air to it will float. It’s also non-flammable which makes it safer than hydrogen.*
 |  |
| 6.16 | Describe the pattern in the physical properties of some noble gases and use this pattern to predict the physical properties of other noble gases* *Boiling point, melting point and density all increase as you go down the group.*
* *See your revision guide for more info*
 | 1d 2c |

#### Topic 7 – Rates of reaction and energy changes Rates of reaction

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| **Students should:** | **Maths skills** |
| *7.1* | *Core Practical: Investigate the effects of changing the conditions of a reaction on the rates of chemical reactions by:*1. *measuring the production of a gas (in the reaction between hydrochloric acid and marble chips)*
2. *observing a colour change (in the reaction between sodium thiosulfate and hydrochloric acid)*
 | 1a, 1c4a, 4b, 4c, 4d, 4e |
| 7.2 | Suggest practical methods for determining the rate of a given reaction* *Using a precipitation reaction to time how low long it takes for an X to disappear*
* *Measure the change in mass*
* *Measure the volume of gas given off*
 | 4b, 4c, 4d, 4ePg 80 ChemPg 129-130 Combined FoundationPg 131 Combined higher  |
| 7.3 | Explain how reactions occur when particles collide and that rates of reaction are increased when the frequency and/or energy of collisions is increased* *More successful collision occur when the particles are moving faster. Particles need to collide with at least the activation energy to be successful.*
 | 1c |
| 7.4 | Explain the effects on rates of reaction of changes in:Temperature – *the hotter it is the faster the rate of reaction. Particles have more energy and speed resulting in more successful collisions.*Concentration – *The higher the concentration, the faster the rate of reaction. There are more particles in the same volume resulting in increased frequency of collisions*Surface area to volume ratio of a solid -  *having a large surface are to volume ratio increases the rate of reactions as the particles around it will have more area to work on so the frequency of collisions increase*Pressure (on reactions involving gases) - *The higher the pressure, the faster the rate of reaction. There are more particles in the same volume resulting in increased frequency of collisions* | 1c, 1d 5c |

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| 7.5 | Interpret graphs of mass, volume or concentration of reactant or product against time*See your revision guide for more info*  | 1c4a, 4d, 4e |
| 7.6 | Describe a catalyst as a substance that speeds up the rate of a reaction without altering the products of the reaction, being itself unchanged chemically and in mass at the end of the reaction |  |
| 7.7 | Explain how the addition of a catalyst increases the rate of a reaction in terms of activation energy* *Catalysts provide an alternate pathway that has a lower activation energy.*
 |  |
| 7.8 | Recall that enzymes are biological catalysts and that enzymes are used in the production of alcoholic drinks* *Enzymes from yeast cells are used in the fermentation process which is used to make alcoholic drinks.*
* *They catalyse the reaction that converts sugars into ethanol and carbon dioxide.*
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#### Heat energy changes in chemical reactions

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| **Students should:** | **Maths skills** |
| 7.9 Recall that changes in heat energy accompany the following changes:a salts dissolving in water b neutralisation reactions c displacement reactions d precipitation reactionsand that, when these reactions take place in solution, temperature changes can be measured to reflect the heat changes |  |
| 7.10 Describe an exothermic change or reaction as one in which heat energy is given out |  |
| 7.11 Describe an endothermic change or reaction as one in which heat energy is taken in |  |
| 7.12 Recall that the breaking of bonds is endothermic and the making of bonds is exothermic |  |
| * 1. Recall that the overall heat energy change for a reaction is:
		1. exothermic if more heat energy is released in forming bonds in the products than is required in breaking bonds in the reactants
		2. endothermic if less heat energy is released in forming bonds in the products than is required in breaking bonds in the reactants
 |  |
| * 1. **Calculate the energy change in a reaction given the energies of bonds (in kJ mol–1)**
* *Overall energy change = energy required to break bonds – energy released by forming bonds*
 | 1a, 1c |
| * 1. Explain the term activation energy
* *The minimum energy that particles need to react when the collide*
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#### Topic 8 – Fuels and Earth science Fuels

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| **Students should:** | **Maths skills** |
| 8.1 | Recall that hydrocarbons are compounds that contain carbon and hydrogen only |  |
| 8.2 | Describe crude oil as:1. a complex mixture of hydrocarbons
2. containing molecules in which carbon atoms are in chains or rings (names, formulae and structures of specific ring molecules not required)
3. an important source of useful substances (fuels and feedstock for the petrochemical industry)
4. a finite resource *(non-renewable)*
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| 8.3 | Describe and explain the separation of crude oil into simpler, more useful mixtures by the process of fractional distillation* *oil is heated. It turns into a gas. The vapours rise up the column. The column has a temperature gradient (it’s hot at the bottom and cold at the top). Longer hydrocarbons with a higher boiling point drain off at the bottom as liquids)*
 |  |
| 8.4 | Recall the names and uses of the following fractions: a gases, used in domestic heating and cooking1. petrol, used as fuel for cars
2. kerosene, used as fuel for aircraft
3. diesel oil, used as fuel for some cars and trains
4. fuel oil, used as fuel for large ships and in some power stations
5. bitumen, used to surface roads and roofs
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| 8.5 | Explain how hydrocarbons in different fractions differ from each other in:1. the number of carbon and hydrogen atoms their molecules contain – *increases as you go down the group*
2. boiling points – *longer chains have stronger intermolecular forces resulting in a higher boiling point*
3. ease of ignition – *shorter hydrocarbons are easier to ignite as they have a lower boiling point*

d viscosity – *longer chains are highly viscous (thick)*and are mostly members of the alkane homologous series | 4a, 4c |

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| 8.6 | Explain an homologous series as a series of compounds which: a have the same general formula1. differ by CH2 in molecular formulae from neighbouring compounds
2. show a gradual variation in physical properties, as exemplified by their boiling points – *the longer the chains, the higher their melting and boiling points*
3. have similar chemical properties
 | 1c, 1d 4a |

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| **Students should:** | **Maths skills** |
| * 1. Describe the complete combustion of hydrocarbon fuels as a reaction in which:
		1. carbon dioxide and water are produced
		2. energy is given out
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| * 1. Explain why the incomplete combustion of hydrocarbons can produce carbon and carbon monoxide
* *There is not enough oxygen*
 |   |
| * 1. Explain how carbon monoxide behaves as a toxic gas
* *Combines with red blood cells instead of oxygen*
 |  |
| 8.10 Describe the problems caused by incomplete combustion producing carbon monoxide and soot in appliances that use carbon compounds as fuels |  |
| 8.11 Explain how impurities in some hydrocarbon fuels result in the production of sulfur dioxide* *Fuels contain sulfur impurities. When they are burned they release CO2*
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| 8.12 Explain some problems associated with acid rain caused when sulfur dioxide dissolves in rain water* *Acid rain acidifies lakes, damages buildings, kills trees and corrodes metals.*
 |  |
| 8.13 Explain why, when fuels are burned in engines, oxygen and nitrogen can react together at high temperatures to produce oxides of nitrogen, which are pollutants |  |
| 8.14 Evaluate the advantages and disadvantages of using hydrogen, rather than petrol, as a fuel in cars*- Advantages = Clean, only produce water. Renewable.**- Disadvantages = Require a special engine which is expensive. Hard to store. Manufactured from fossil fuels.* |  |
| 8.15 Recall that petrol, kerosene and diesel oil are non-renewable fossil fuels obtained from crude oil and methane is a non- renewable fossil fuel found in natural gas |  |
| 8.16 Explain how cracking involves the breaking down of larger, saturated hydrocarbon molecules (alkanes) into smaller, more useful ones, some of which are unsaturated (alkenes)- *A vapourised hydrocarbon is passed over a powdered aluminium oxide catalyst at 400-700°C and 70atm.* | 1c |
| 8.17 Explain why cracking is necessary*- there is more of a demand for smaller chains like petrol and diesel.**- cracking also produces alkenes that can be used to make polymers* | 2c |

#### Earth and atmospheric science

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| **Students should:** | **Maths skills** |
| 8.18 Recall that the gases produced by volcanic activity formed the Earth’s early atmosphere |  |
| * 1. Describe that the Earth’s early atmosphere was thought to contain:
		1. little or no oxygen
		2. a large amount of carbon dioxide

c water vapourd small amounts of other gasesand interpret evidence relating to this | 2c 3a 4a |
| 8.20 Explain how condensation of water vapour formed oceans |  |
| 8.21 Explain how the amount of carbon dioxide in the atmosphere was decreased when carbon dioxide dissolved as the oceans formed |  |
| 8.22 Explain how the growth of primitive plants used carbon dioxide and released oxygen by photosynthesis and consequently the amount of oxygen in the atmosphere gradually increased |  |
| 8.23 Describe the chemical test for oxygen*- relights a glowing splint* |  |
| 8.24 Describe how various gases in the atmosphere, including carbon dioxide, methane and water vapour, absorb heat radiated from the Earth, subsequently releasing energy which keeps the Earth warm: this is known as the greenhouse effect |  |
| * 1. Evaluate the evidence for human activity causing climate change, considering:
		1. the correlation between the change in atmospheric carbon dioxide concentration, the consumption of fossil fuels and temperature change
		2. the uncertainties caused by the location where these measurements are taken and historical accuracy
* historically less data was taken over fewer locations. If you go back far enough there are no records at all.
 | 2c, 2h 4a |
| * 1. Describe:

 a the composition of today’s atmosphereb the potential effects on the climate of increased levels of carbon dioxide and methane generated by human activity, including burning fossil fuels and livestock farmingc that these effects may be mitigated: consider scale, risk and environmental implications |  |